

Errata for the Second Printing of Discovering Design with Chemistry

**Student Text**

- p. 196:** For 16 (d), the reaction should be  $3\text{Mg (s)} + \text{N}_2 \text{(g)} \rightarrow \text{Mg}_3\text{N}_2 \text{(s)}$
- p. 240:** In the last paragraph, the addition of the two masses should be “23.4 g + 4.91 g = 28.3 g.” That means “28.31” should be changed to “28.3” the three times it is found on the page.
- p. 351:** Two lines above the last equation on the page, “NaOH” should be changed to “H<sub>3</sub>PO<sub>4</sub>.”
- p. 351:** The NaOH should be changed to LiOH
- p. 384:** The salt bridge should be labelled.
- p. 451:** In #4, there is a mismatch between the units in the question and its answer. Both units should just be kJ.
- p. 451:** In #9, it should ask for the  $\Delta H_f^\circ$  of H<sub>2</sub>S (g), not H<sub>2</sub>S (l).

**Answer Key**

- p. 33:** For problem 9, “highest” should be replaced with “higher” in the underlined text.
- p. 66:** The phase symbol “(g)” should follow each molecule in first chemical equation.
- p. 176:** There is a mismatch in units between the question and the answer for #4. The unit should be just kJ in both.
- p. 221:** Top of the page should read:

$$\Delta H = (1 \text{ mole}) \cdot \left(0 \frac{\text{kJ}}{\text{mole}}\right) + (3 \text{ moles}) \cdot \left(0 \frac{\text{kJ}}{\text{mole}}\right) - (2 \text{ moles}) \cdot \left(0 \frac{\text{kJ}}{\text{mole}}\right)$$

This doesn't affect the answer.